Survey of Chemistry II (CHEM1152) Syllabus- Fall 2016- GEORGIA STATE UNIVERISTY, Chemistry Department

As on August 16/2016… If changes are needed, changes will be made. Do not make plans on school time as there are no make ups in this course. Welcome to CHEM1152-

Instructor: Dr. Angela Maria Navarro-Eisenstein-

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Office hours: Tuesdays, Thursdays 9:00-10:00 or by appointment (Never on Fridays- too many meetings)

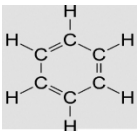
Prerequisite: Successful completion of Chemistry 1151K You need to know these topics before you start organic Chemistry

August 16- August-21 Review topics from prerequisite 1151: **What do you need to review from CHEM1151? I will not teach these topics in our summer class: Here is a summary**

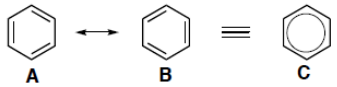
- ✓ Review density from **chapter 1** for lab experiment.
- ✓ Review molarity from **chapter 8** for lab experiment.
- ✓ Review acids and bases concepts: stomach acid, neutralization reaction between acid and base, chapter 10 (none of pH calculation).

Study chapter 5 to refresh the following topics.

- ✓ The octet rule, atoms will lose, gain or share their outermost valence electrons to get 8 electrons around central atom, in order to fulfill stability of a noble gas.
- ✓ Metals from Group I-A easily lose one electron becoming cations (positively charged) Na+, K+,
- ✓ Nonmetal from Group VIIA such as F, Cl, Br easily gain one electron to become anions (negatively charged) (Cl-, Br-).
- ✓ Electronegativity and bond polarity, F, O, N, Cl, Br all pull electron density to the nucleus when forming a polar covalent bond; those are the most electronegative atoms (all from the top right corner of the periodic chart).
- ✓ Chemical covalent bonds: Two nonmetals sharing electrons evenly, (non polar covalent bond, example: Cl-Cl for chlorine gas (Cl_2 element) or two nonmetals sharing electron unevenly (polar covalent bond) like H-F, since fluorine is more electronegative than hydrogen, there is more electron density
- ✓ Lewis electron-dot structures and bonds types for H, C, N, O, S, P and halogens. For Lewis representations, follow the old ACS periodic chart. Elements from Group I-A carry one electron, one dot H·, Li·, Na·, K· because their electron distribution ends as S^1, which means H shares one electron (nonmetal). All other elements on group IA transfer one electron, becoming cations.
- ✓ Explore electron-dot structures: C, 4 electrons, N five, O six, F, Cl, Br, seven. Recall only 2 electrons can be around each side of the box.
- ✓ Bond capacity: Hydrogen or halogen always make one bond, like H-Cl C, 4 bonds and Phosphorous 5 bonds,
- ✓ Bond order: Carbon always makes 4 bonds (single, C-C, double C=C or triple C≡ C)
- ✓ Nitrogen always makes three bonds, it can be single double or triple (and it has one set of lone pairs of electrons)
- ✓ Oxygen always makes two bonds, it can be two single bonds or one double bond (and it has two sets of lone pairs of electrons).
- ✓ Resonance structures, this is when electron delocalization takes place by using empty p orbitals, sideways overlap (no head to head). When exploring resonance structures atoms don't move, only electron pairs are delocalized. For example O_3 and benzene C_6H_6 have electrons that can be placed in multiple positions to fulfill the octet of all atoms; that implies there are different representations or structures for the same exact compound. They are called resonance contributors.



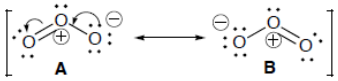
C_6H_6



2 Resonance Structures

Example: O_3 (ozone)

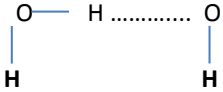
O_3



2 Resonance Forms

- ✓ Attractive forces in compounds: London Dispersion forces (between non polar species-similar polarity substances easily intermingle with one another- for example olive oil easily mixes with Vaseline- petroleum jelly, because they have same nature- lots of C-H), however, olive oil (nonpolar) does not mix with water (polar substance) London Dispersion forces same type as Van der Waals interactions.
- ✓ Ionic attraction (salts): When two balanced ions combine they form a salt (a crystal- like NaCl, table salt). In the organic part you will see organic acids form polyatomic anion like carbonate $HCOO^-$
- ✓ Hydrogen bonding Hydrogen atom partial positive partial charge is shared between two electronegative atoms,
- ✓ It could be N-H…….O=, or –O-H…….O-H or N-H……N

Example water and water



H **H** **Where the intermolecular H bond force is represented by ……..**

What is new? Visit D2L/Brightspace/Icollege. Introduction to organic chemistry is a video and alkanes' discussion will be in the classroom. Read the chapter. After each lecture make a habit of the following: Finish worksheets provided or you print from my resources, Watch IPad videos, re-read the chapter, go to Mastering Chemistry. **Please place a copy of syllabus in a three rings 2" binder. Print documents and keep them organized.**

This course comprises of two parts:
First part is **organic chemistry** where you will have 4 homework, 2 tests and one ACS final.
First part is **biochemistry** where you will have 4 homework, 2 tests and one ACS final.
This is a hybrid course: I expect you watch all videos for Alkanes before first day of class: Chapter 11 Alkanes and cycloalkanes, you need to know alkyl groups for the first week of class-
Methyl (CH_3-, Ethyl CH_3CH_2-, Propyl, CH_3CH_2 CH_3 Butyl $CH_3CH_2CH_2CH_2$ by heart to name branched compounds
Pay close attention to isopropyl, and isobutyl as their prefixes are used in the alphabetization of the names of organic compounds. Nomenclature of alkanes of complex molecules is a video and will not be discussed in lecture.

Textbook: Karen C. Timberlake **General, Organic, and Biological Chemistry: Structures of Life (4th Edition)** (Prentice Hall, PEARSON) ISBN, 13-978-0-321-75089-1, ISBN, 10-0321-75089-6 If you have taken CHEM1151 during the Fall 2014, or Spring-Summer 2015 or Fall 2015 you already have the book. If you have any other General Organic Biochemistry Book, you can use it, just follow the order of topics listed. It makes no difference what author, edition you get as I will provide lots my own resources. Buy the least expensive book you find in Amazon®.

Laboratory Manual: A Laboratory Manual for Chemistry1152K will be available to students during the first pre-laboratory lecture at the scheduled time for each student to attend and check-in. **There is not lab during the first week of semester neither the holiday (Labor Day or MLK).** Class and lab resume after holidays. Don't take vacation during this time as that will not be a valid absence with excuse

When sending an e-mail to Dr. Navarro-Eisenstein, **under the subject heading**, please state in which class you are enrolled. (i.e., CHEM 1050, CHEM 1152, CHEM 1211, CHEM 1212, CHEM 2400, CHEM 3100 CHEM 3110, CHEM 4600, CHEM 6600)- Chemistry is not enough information. It is important you email to anavarro@gsu.edu from your GSU account to receive a prompt response. If you email Dr. Navarro from D2L/Brightspace/Icollege, I will only reply your email when I access to enter grades. Please, do not e-mail through Gmail, Hotmail, Yahooh etc. as your message will go to span. I will email the class from GSU mass collective emails. If the subject does not apply to you, please ignore it.

Storms: In the event of a storm students are expected to review topics discussed on line videos as all material will be included in assessments. Students will have to watch the videos and read the book and slides understanding that extra instruction or clarification will only be available during office hours. **In the event of cancelation due to storms an updated schedule of assessments will be posted.** Topics discussed in videos are supplemental instruction and will not be discussed in detail in the classroom.

Office Hours: 434-C KELL HALL are "First come first serves", if a large group comes all will be seen together if student coming first does not bring personal –private questions.

Use the office hours: I will help you if you approach me with your questions and /or concerns. **Do not ask me at the end of semester "What do I need to do to get an A". Do not ask me to allow you to write a paper to raise your grade. I will give same opportunities to all students. This is doable.**

Attend SI. Our leader has the experience from taking this course. She will give you hints and tell you what works and what doesn't in a course with so many topics like this. Besides she will guide you when working problems. Understand she will not do your homework; that would hurt you as you and only you take exams in class.
Grading: Lecture comprises 75% of the overall course grade and laboratory 25%. GSU does not provide two separate grades in the transcript.

The lecture grade will be computed from homework, six major tests, and two final examinations. **The final exams are standardized American Chemical Society comprehensive exams. Every topic and chapter discussed will be in the ACS finals. The final exams** will count as 1/4 of the total lecture grade. These standardized exams consist of a total of 120 multiple choice questions. Students will have 55 minutes to complete 60 multiple choices each part - 60 questions in the

Organic part (MIDTERM) and 60 questions in the Biochemistry (FINAL). These exams scores will be curved based on national percentiles using national norms from ACS.

Grades Tentative Cut Off:
A+*: 96 % A*: 90%; A-: 87%; B+: 84% B: 80% B-: 77%, C+: 74% C: 70% C-: 67%, D= 64-66%, F= below 64%

***To get an A or A+ in the course students have to get at least 86 percentile in the ACS exams. (40-44/60 questions depending on the test provided by the Chemistry Department)**

Letter grade percentage is = total number of points accumulated x 100
800

200 points	**laboratory portion includes a final exam**
200 points	**2 ACS finals @ 100 pts each**
130 points	**Homework on line and hard copy**
135 points	**3/organic in class tests, 35 minutes lecture after tests**
135 points	**3/ biochemistry in class tests, 35 minutes lecture after tests**
Total	**800 points**

Preparation for the course: Read the chapter to be discussed before you come to lecture. Work the problems within the chapter, as they bring step by step how to arrive to the answers. Work the problems at the end of the chapters. You know what you know when you answer the questions with the closed book. Exams are closed book. Previewing solutions to problems gives a false sense of confidence about the subject matter, and typically results in poorer test scores.
Planning ahead is a key to success. Your performance in science is a lot better when you study daily. Do not wait until the night before the exam to begin studying. You EARN your grade, I DO NOT GIVE grades. As you read the material, you must take written notes and **underline**. Use **highlighters** or color pens. That will help you throughout the semester and to study for finals.

Miscellaneous:

- Make-ups: There is no chance for make ups due to time constraint. Plan accordingly.
- Academic Honesty: The honor code embraced by universities expresses an ideal of character, conduct, and citizenship. This applies especially to academic honesty and integrity. Passing off someone else's work as your own represents intellectual fraud and theft, and violates the core values of academic community. You will be asked to sign up the honor code in bold as written beneath, along with your printed name on the first page of assessments. "**As a member of the student body taking this course, I consider myself bound, guaranteed and compelled by honor to develop and uphold high standards of honesty and behavior."**
- Class Preparation: Suggested reading assignments should be completed before the next lecture.
- Attendance: Students are expected to attend all lectures. As a courtesy to your fellow students, please arrive on time and do not leave during the lecture. **Please bring me a schedule of your RELIGIOUS HOLIDAYS OBSERVANCE the SECOND WEEK of class.**
- **Disabilities: Students who wish to request accommodation for a disability** may do so by registering with the Office of Disability Services. Students may only be accommodated upon issuance by the Office of Disability Services of a signed.
- Accommodation Plan: Students are responsible for providing a copy of that accommodation plan to instructors of all classes in which an accommodation is sought. Your need for accommodations will only be discussed in the office and never at the beginning or end of class **and never in front of classmates.**
- Electronic devices: Students are requested not to bring cellular telephones and/or Ipads or tablets to exams. Cell phones and laptops are allowed during lectures, BUT NOT during quizzes or exams
- The University requires that faculty members must, on a date after the mid-point of the course to be set by the Provost (or his designee): A) **Give a WF** to all those students who are on their rolls but no longer taking the class and B) **Report the last day** the student attended or turned in an assignment.

- Students need to **show their GSU Panther ID card when taking any test, quiz or exam**. All tests/exams are taken in class. The basic ideas and principles on these exams come from the book and lecture material and are designed to test a student's 1) understanding of the concepts and 2) ability to solve problems, as well as 3) knowledge of the facts.
- Research shows that the more **different ways** you present information to the brain the easier it is to learn. In other words **hear it, see it, say it, write it, practice it, highlight it, quiz it,** etc.
- In general, questions and problems on exams will be original and not copied from those found in a chapter of the textbook or old previous exams given. Don't waste time memorizing answers from old exams as the instructor modifies questions every semester. Exams are designed so that the majority of questions are of medium difficulty, some are relatively easy and very few are challenging.
- Major tests and the final exams will be multiple choices (green scantrons). Missing any exam will result in the assignment of a zero for that exam, in either the lecture or the laboratory. **NO MAKEUP EXAMS ARE GIVEN. Exams cannot be taken earlier or later than day scheduled.** Students missing an exam will be expected to submit a written note explaining why the exam was missed, to provide valid evidence for that excuse and to discuss absence in the office, no in front of the class.
- Any student presenting falsified documentation will receive an "F" for the course and be referred to the Chemistry Department Chair or Dean of Students for disciplinary action.
- The lecture professor may retain copies of students' homework, quizzes or tests. Students can retrieve graded assessments during office hours. Students are responsible for keeping all tests and other papers until after they have received their final grade for this course.
- Two (2) blue large Scantrons for ACS the Final Exams **Form no. 4521** (blue form used by the Testing Center's machine) will be needed for this course. They can be obtained from at the bookstore or for free in the Students Center. **If the Chem. Dept. has the blue scantrons, I will announce in class.** Those scantrons are graded in the Testing and Counseling Center, Piedmont 75.
- **Six (6) green** small scantrons are required for in class tests for Chem. Dept. machine). I do not return original graded scantrons, but students are welcome to look at them during office hours. **We will have lecture after tests.**
- **Class Preparation: The price of success** is high. Anything of value requires great effort. You have to work hard, be persistent, and pay attention to details. These traits are ultimately why a college degree is valuable, plus the capacity to learn. **Believe you can succeed.** Be willing to pay the price. **Accept responsibility for your learning!** It is your choice. I could be the best teacher in the world but your performance depends only on the time and effort you invest in this course. Chemistry is a highly structured subject in which each new topic is based on others previously discussed. Therefore, if one topic is not mastered, it becomes *increasingly* difficult to master those that follow. **Missing even one class can lead to problems** that the average student cannot overcome. Also, chemistry does not lend itself to "cramming". What you learned from the first chapter is needed for the second; what you learned from the third chapter will be needed the last day. Complex concepts build up from beginning chapters. **Attend all lectures!** The quizzes and exams are based mostly on material that is covered in class. You must be present to know what is going on.
- Students are strongly encouraged to download lecture notes from WebCT/Vista/D2L/ Brightspace/**ICollege** before coming to class**.** The lecture **visual aids** for the instructor and **are not intended** to be *"the only source of study"* for the students. **You need to study from the textbook for all exams.**
- **Three habits** will help in mastering each topic as it arises, and will reinforce the topics previously covered:
  - ✓ **Read** the assigned material *before* it is covered in lecture and watch my own Ipad videos. You can explore more in your own time but you might encounter the same topics with more details for higher level chemistry.
  - ✓ **Work** through the **example** and **practice problems** from the textbook within each assigned chapter or those I place in D2L/Brightspace.
  - ✓ **Work** a large number and wide variety of problems "as many **end-of-chapter problems as possible".** All in-chapter examples and practice exercises should be done. To reward your hard work I might use some of the problems from your book for exams and quizzes.

- The instructor may assign or re-arrange seating at any examination or quiz. Please leave first row empty during all examinations.
- The instructor may take up an examination or quiz from any student who is behaving in such a manner as to disrupt the class during the examination or quiz, and assign a grade of zero to that student for that exercise. Such disruptive behavior includes looking around oneself or talking. Even if you are not copying, any form of looking around, neck

stretching, gesturing or talking is unacceptable /considered inappropriate testing procedure, and will result in a zero score for the exam plus subsequent disciplinary action. **Hats need to be removed or turned around.** Students who feel they must move around during these times can apply for special testing privileges through the GSU Department of Disability Services.

- The Instructor reserves the right to move students during the tests. During all exams, sit up straight and keep your paper directly in front of you and out of sight from other students.
- Before an exam is given to the students, there should be nothing on your desk except pencils, erasers, a scantron, and ID (scratch paper will be provided if necessary). Students are required to show (and leave) their student identification on the desk in order to take the test. Tests will be graded ONLY if a student I.D. is shown. Write your name, student ID number and exam color/ or exam letter (A or B or C) on your scantron and set your ID out to be checked after completion of the tests when you and the scantron to the instructor.
- Answer sheets, scantron and exam papers should be completely covered throughout all assessments. If your paper or scantron can be seen, then you are a participant in cheating and **ALL participants will face disciplinary action**, which may include a failing grade for the entire course.
- You **cannot use your cell phone as a calculator or as a watch.** If you need to be on-call during an exam, you can turn your cell phone to "vibrate" and leave it up front with the instructor.
- Use the restroom and complete any other personal business **before coming to an exam.** Bring tissues in case you need them. Students may leave the room only after their exam has been turned in. Leaving the room during any testing procedure will result in a score of zero for the exam.
- Students arriving late to exams will not have any additional time. Make sure you take into consideration Downtown's heavy traffic. Students arriving to ACS finals after instructions are given will not take such final.
- **Cancellation of Classes:** Official closure of the university is determined by the university administration. This sometimes occurs due to inclement weather, in which case notification will be make through local radio and television outlets.
- Should closure result in cancellation of classes or examination periods, resumption of the missed activities would occur at the next regular class period when the university reopens, or as determined by the course instructor. Should an instructor be unable to meet a class for reasons others than those above, another instructor would normally meet the class as scheduled. Be hereby advised, however, that on rare occasions conditions could require the cancellation of class or examination periods. In such cases, there would be *official* notification of the students affected via email, text message and the local news. Should this notification be through notices posted in the classroom or other means, the *student* has the final responsibility of confirming the authenticity of the cancellation through the Chemistry Department Office.
- **Each student has the responsibility of checking their email and D2L/Brightspace/ICollege on a daily basis.**
- **Prohibited Accessories**: Students may **NOT** use a **cell phone** as a timepiece or calculators during exams, or any such transmitting equipment (e.g. **Bluetooth, MP3, laptop, Ipod, Iphone, Ipad).** They are all strictly forbidden during any test or exam. Electronic calculators are not needed in this course, only for the lab portion.
- **Classroom etiquette:** Please do not reserve a seat for your friends coming late. Please **refrain from chit-chatting** during class as it distracts not only the students who sit around you but also the instructor. Disruptive conduct during class will not be tolerated and **appropriate action will be taken against you** (refer to your copy of the Student Conduct Code).
- **Cheating:** All tests taken must represent your individual, unaided efforts. To receive or offer information during an examination is cheating. The use of unauthorized supplementary materials during tests is also cheating. A student who cheats on an exam will receive **a zero for that exam which cannot be dropped as the lowest grade**. Any suspected offenses may also be referred to the Department Chairman and/or the Dean of Students for appropriated disciplinary action. The Department of Chemistry follows the university policy on academic honesty published in the "Faculty Affairs Handbook" and the "On Campus: The Undergraduate Co-Curricular Affairs Handbook."

Georgia State University Student Conduct and Integrity Policy: *The Georgia State University Policy on Academic Honesty* is applicable to this course, including but not necessarily limited to infractions in the areas of ***plagiarism,*** *cheating on examinations, unauthorized collaboration, falsification,* and *multiple submissions*. This policy is published in *On Campus: the Student Handbook*, available to all members of the university community. **Also all results submitted by you in the laboratory section of this course must reflect your individual effort**. Only original data obtained by your experimentation in the assigned laboratory may be used, except when specifically authorized by your laboratory instructor. Supplementary source material (handbooks, reference literature, etc.) must be *clearly referenced* (title, author, volume, page(s), etc.).

Falsification or destruction of data constitutes cheating. **Conduct or actions that disrupt class,** examination, or laboratory periods or falsification of information related to chemistry courses by any student will be taken as a violation of the policies of the Board of Regents of the University System of Georgia and the Georgia State University Student Code of Conduct, Section 6.0. Any suspected offenses may be referred to the Department Chair.

Some Examples of Unacceptable Student Conduct:

- ✓ Leaving class before the lecture is over
- ✓ Leaving class after taking a quiz or turning a homework
- ✓ Talking while your professor is lecturing distracting your fellow classmates
- ✓ Arguing with the professor about student conduct
- ✓ No following the testing procedures as listed in this syllabus
- ✓ No sitting up straight with paper directly in front of you during a quiz or exam or no keeping scantron covered
- ✓ No having your student ID for a quiz or test
- ✓ Letting your cell phone ring audibly during a lecture or exam or having a cell phone available during a quiz or test
- ✓ Using a disrespectful tone of voice, harsh words
- ✓ Using profanity or making inappropriate gestures of any kind

Profile of the A/B student after Dr. Terry L. Fry

Note: This checklist might be used by students to determine why they are or are not attaining their desired goals in class. Past students have contributed items to this list which they believed were characteristic of A/B students

Attends to school regularly
It is not tardy to class
Always writes down homework assignments
Does homework fully and completely (neatness counts)
Does make up work promptly after an absence
Is well organized
Schedules a regular study time
Comes to class prepared (books, paper, pencil, homework)
Pays attention in class (does not sleep or goof off)
Asks questions in class
Participates in class discussion
Develops good concentration (don't study with music or TV on)
Is courteous to others (does not distract students and instructor with cell phone or lab top)
Takes good lecture notes
Does not allow problems unrelated to school affect school work
Does extra work without expecting a grade
Sets and achieve goals and cooperates with the instructor

APPENDIX

TENTATIVE SCHEDULE OF ACTIVITIES: **The schedule below is only tentative and it is subject to changes if needed**

In order to get familiar with the new vocabulary, the students are required to read the chapters before the topics are discussed and do THE TEXTBOOK practice solved problems within the chapters. After discussions, work those at the end of chapters. Students are required to check announcements of changes in D2L – Brightspace and to attend to lecture as this is only a general plan for the course; deviations may be necessary. Important: In the event of a storm students are expected to review topics discussed on line as all material will be included in assessments. In the event of cancelation due to storms an updated schedule of assessments will be posted. Work problems within the chapter. Make flash cards for functional groups. Finish worksheet, and work homework at the tutoring center (Courtland North 217) if needed or use your book or Watch IPad videos.

Tentative Topic and Reading Assignment Before Each Lecture

Date	Tentative Topic and Reading assignment before each lecture The first day of class (make flashcards for functional groups)
	Part I – ORGANIC CHEMISTRY
Tuesday 8/23	Chapter 11 Introduction to organic chemistry and alkanes
Thursday 8/25	Chapter 11 Alkanes and cycloalkanes,
Tuesday 8/30	Chapter 11 Alkanes and cycloalkanes (homework 1 due)
Thursday 9/1	Chapter 12 Alkenes, alkynes
9/5	Labor day –no class-no labs (make flashcards for benzenes)
Tuesday 9/6	Test # 1; chapter 11, part of 12, (40 POINTS) 35 min.Lecture- Chapter 12 Cycloalkenes, alkynes
Thursday 9/8	Chapter 12 Cycloalkenes, -continued, Benzene and aromatic compds. Chapter 13 Alcohols
Tuesday 9/13	Chapter 13 Alcohols/(homework 2 due)
Thursday 9/15	Chapter 13 Phenols, ethers, thiols, sulfides, disulfides
Tuesday 9/20	Test #2 (40 POINTS) includes chapters 12,13 35 min. Lecture- Chapter 18 Amines
Thursday 9/22	Chapter 18 Amines/ Chapter 14, Aldehydes, ketones Chapter 15 carbohydrates- introduction, stereochemistry and mono saccharides (glucose, and fructose, ribose)
Tuesday 9/27	Chapter 14, Aldehydes/ hemiacetals/acetals-intro to carbs/stereochemistry (homework 3 due)
Thursday 9/29	Chapter 16, Carboxylic acids/ intro to aminoacids/intro to lipids
Tuesday 10/4	Chapter 16 (Carboxylic acids derivatives-(esters) and 18 (amides)
Thursday 10/6	Test #3 (55 POINTS) includes 13-reactions, 14, 16, 18 35 min. Lecture- Hydrolysis (homework-4 due)
Tuesday 10/11	**MIDPOINT: This is the last day to withdraw from class & receive a "W".**
	Part II - BIOCHEMISTRY
Tuesday 10/11	Midterm-ACS FINAL EXAM (ORGANIC) (100 POINTS) cumulative 11,12,13,14,16,18
Thursday 10/13	Chapt/15 carbohydrates (make flashcards for Amino acids)
Tuesday 10/18	Chapt/15 carbohydrates if needed/Chapter 17 Lipids
Thursday 10/20	Chapter 17 Lipids (homework 5* due)
Tuesday 10/25	**Test #4 (40 POINTS) includes chapters 15,17** 35 min. Lecture- **-Chapter 19 amino acids**
Thursday 10/27	Chapter 19 proteins / and Chapt. 20 enzymes
Tuesday 11/1	Chapter 20 enzymes and vitamins
Thursday 11/3	Chapter 21 DNA, RNA (homework 6* due)
Tuesday 11/8	**Test #5 (40 POINTS) includes chapters 15,17 and part of 21** 35 min. Lecture- Chapter 21 DNA, RNA
Thursday 11/10	Chapt. 22 Metabolism,(NAD+, FAD, AcCoA) glycolysis gluconeogenesis, Glycogenolysis
Tuesday 11/15	Chapt. 24, beta-oxidation) lipids metabolism
Thursday 11/17	Chapter 23 citric acid cycle (Krebs Cycle), Electron transport chain (ETC)
No class	Thanksgiving break Nov-21 to Nov 27
Tuesday 11/29	Chapter 23 ETC and ATP synthesis "oxidative phosphorylation"
Thursday 12/1	Test #6 (55 POINTS) (Homework 7* due
Tuesday 12/5	No class- reading days
Thursday 12/8	Last day of classes ACS FINAL EXAM (BIOCHEMISTRY) (100 POINTS) at 10:45 am

CHEM1152K-1212K-lab schedule Class Fall schedule

Mondays	***Tuesdays***	***Wednesdays***	***Thursdays***
Jie Jiang 1152K L **CRN**84751 lab lecture 8:00-8:45 N218/ lab 8:55-10:30 /N348	B. Canup 1152K L CRN 82160 lab lec 8:00-8:45 N218 /lab 8:55-10:30 /N348		
	Office hours at 9:00-10:00 AM For 1152	***Office hours at 9:00-10:00 AM For 1212***	***Office hours at 9:00-10:00 AM For 1152***
Navarro **1212KL** **CRN-80131 (Bar)** **10:30-11:45 lab lec P362/** **/lab 11:55-2:30 P355** **(12/5)**			
Office hours 3:00-4:00pm For 1212	**Navarro** **1152K-177 (178)** **CRN 84751, 83785, 82160 -TR** **11:00-12:15** **ADHOLD 5**		**Navarro** **1152K-177 (178)** **CRN 84751, 83785, 82160 -TR** **11:00-12:15** **ADHOLD 5**
	Navarro CRN 84635 (Fin) lab lec 1:30-2:45 P362 /lab 2:45-5:30 P355 (11/29)		Abdelwahab 1152K L **CRN** 83785 lab lecture 2-2:50 N 218 /lab 3:00-4:30 N348

During the spring of 2016 many students came for advisement and I realized lack of organization was hurting them. You need to have the discipline of writing down everything you do in a schedule and follow it. If I don't do it I would miss important deadlines in my career. Old school works better than the new apps in cases like this. Print it and put it where you can see it daily. CEOs don't go to sleep late. They get up early. The brain can process double the information given after a good night of sleep. You get up and have a good healthy breakfast and topics you read or videos you watch will be retained before you bombard the day with other things. This summer give it a shot, get up at 5:00 am-5:30 am every day to read and tell me later the outcome.

My OWN school schedule

Time	Monday	Tuesday	Wednesday	Thursday	Friday
5:00 AM	Get up	Get up	Get up	Get up	Get up
5:30 AM	Breakfast/shower				
5:45AM	Read the chapter of today	Read the chapter of today	Read the chapter of today	Read the chapter of today	Read the chapter of today
7:00 AM					
8:00AM					
9:00 AM					
10:00 AM					
11:00AM					
12:00 noon					
1:00 PM					
2:00 PM					
3:00 PM					
4:00 PM					
5:00 PM					
6:00 PM					
7:00 PM					
8:00 PM					
9:00 PM					
10:00 PM	Go to bed, even if I am not sleepy	Good sleep is needed for your health	If I study daily, I will get an A in this course	A brain that rests will get better scores in all tests.	I can do whatever I want tonight. There is no school tomorrow.